Common Polyatomic Ions

Name(s)	Formula	Name(s)	Formula
ammonium	NH_4^+	iodate	IO_3^-
acetate	CH_3COO^- $C_2H_3O_2^-$	nitrate	NO_3^-
bromate	BrO_3^-	nitrite	NO_2^-
carbonate	CO_3^{2-}	oxalate	$C_2O_4^{2-}$
chlorate	ClO_3^-	perchlorate	ClO_4^-
chlorite	ClO_2^-	periodate	IO_4^-
chromate	CrO_4^{2-}	permanganate	MnO_4^-
cyanide	CN^-	peroxide	O_2^{2-}
dichromate	$Cr_2O_7^{2-}$	phosphate	PO_4^{3-}
hydrogen carbonate bicarbonate	HCO_3^-	phosphite	PO_3^{3-}
		silicate	SiO_4^{4-}
hydrogen sulfate bisulfate	HSO_4		
		sulfate	SO_4^{2-}
hydrogen phosphate biphosphate	HPO_4^{2-}	sulfite	SO_3^{2-}
		thiocyanate	SCN^-
hydroxide	OH^-	thiosulfate	$S_2O_3^{2-}$
hypochlorite	ClO^-		

Metal Ions with more than one oxidation number

copper (I)	cuprous	Cu^+
copper (II)	cupric	Cu^{2+}
iron (II)	ferrous	Fe^{2+}
iron (III)	ferric	Fe^{3+}
lead (II)	plumbous	Pb^{2+}
lead (IV)	plumbic	Pb^{4+}
mercury (I)	mercurous	Hg^+
mercury (II)	mercuric	Hg^{2+}
tin (II)	stannous	Sn^{2+}
tin (IV)	stannic	Sn^{4+}